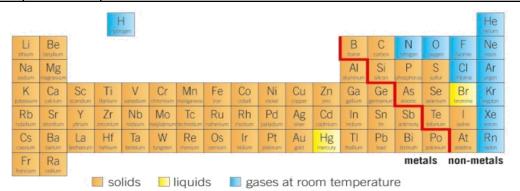
## **Chemistry 2.1: The Periodic Table**

## **Section 1: The Periodic Table**

	2 Metals	letals Found on the left side whereas non-metals are on the right.	
:	Metalloids	Near the <b>stepped line</b> and share properties of both metals/non-metals.	
Ĺ	The vertical columns found on the <b>Periodic Table</b> . These elements share similar properties such as <b>density</b>		
5	Period	The horizontal rows found on the <b>Periodic Table.</b>	



Section 2: Metals and non-metals					
	Properties of metals and non-metals	Metal	Non-metal		
6	Good <b>conductor</b> of electricity	Υ	N		
7	Good <b>conductor</b> of heat	Υ	N		
8	Appearance?	Shiny	Dull		
9	Density? (Mass of a material in a certain volume)	High	Low		
10	Malleable (can be hammered into shapes)	Υ	Breaks		
11	Ductile? (can be pulled into a wire)	Y	Breaks		
12	Sonorous? (makes a ringing sound when hit)	Υ	N		

Se	Section 3: Elements of Group 1			
13	The Alkali Metals	Lithium (Li), Sodium (Na), Potassium (K), Rubidium (Rb), Caesium (Cs) and Francium (Fr).		
14 Reactive The ability to take part in chemical reactions.  Reactivity of Group 1  The ability to take part in chemical reactions.  As we move down group 1, the elements become more reactions.		The ability to take part in <b>chemical reactions</b> .		
		As we move down <b>group 1</b> , the elements become <b>more</b> reactive		
ı	These are the word	Lithium + Water → Lithium hydroxide + hydrogen		
	<b>equations</b> for group 1 metals	Sodium + Water → Sodium hydroxide + hydrogen		
	reacting with water:	Potassium + Water → Potassium hydroxide + hydrogen		















## Section 4: Elements of Group 7

17	The Halogens	Fluorine (F), Chlorine (Cl), Bromine (Br), Iodine (I) and Astatine (At).
18	Reactivity of group 7	Fluorine most reactive: As we move down <b>group 7,</b> the elements become <b>less</b> reactive.
19		A more reactive element pushes out a less reactive element from its compound in a chemical reaction.
20		Elements nearer the top of <b>group 7 displace</b> elements lower in the group: fluorine will displace chlorine; chlorine will displace bromine
21	Example of <b>displacement</b>	Chlorine + potassium bromide → potassium chloride + bromine

Section 5: Elements of Group 0				
22	The Noble Gases	Helium (He), Neon (Ne), Argon (Ar), Krypton (Kr), Xenon (Xe) and Radon (Rn).		
23	Reactivity of group 0	The elements in <b>Group 0</b> are <b>unreactive.</b>		
24	Melting and boiling points	The <b>melting/boiling point</b> for the <b>Noble gases increase</b> as you move down the group.		